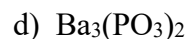
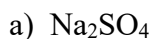


### Practice Naming Compounds

1. Which of the following are ionic compounds? Which are covalent compounds? Name each compound.

#	Formula	Ionic or Covalent?	NAME
a)	MgF <sub>2</sub>		
b)	CCl <sub>4</sub>		
c)	NaC <sub>2</sub> H <sub>3</sub> O <sub>2</sub>		
d)	Fe(NO <sub>3</sub> ) <sub>3</sub>		
e)	CoPO <sub>4</sub>		
f)	Al <sub>2</sub> S <sub>3</sub>		
g)	HCl		
h)	(NH <sub>4</sub> ) <sub>2</sub> CO <sub>3</sub>		

2. Name the following ionic compounds:



3. Give the formulas for the following ionic compounds:

a) aluminum hydroxide

b) iron (III) nitride

c) magnesium phosphate

d) ammonium sulfate

e) silver carbonate

f) manganese (II) nitrate

4. Give the formulas for the following covalent compounds:

a) sulfur trioxide

b) hydrobromic acid

c) carbonic acid

d) dinitrogen tetrachloride

### Answers to Chapter 4 Practice Questions

1. Which of the following are ionic compounds? Which are covalent compounds? Name each compound.

#	Formula	Ionic or Covalent?	NAME
a)	MgF <sub>2</sub>	I	magnesium fluoride
b)	CCl <sub>4</sub>	C	carbon tetrachloride
c)	NaC <sub>2</sub> H <sub>3</sub> O <sub>2</sub>	I	sodium acetate
d)	Fe(NO <sub>3</sub> ) <sub>3</sub>	I	iron (III) nitrate
e)	CoPO <sub>4</sub>	I	cobalt (III) phosphate
f)	Al <sub>2</sub> S <sub>3</sub>	I	aluminum sulfide
g)	HCl	C	hydrochloric acid
h)	(NH <sub>4</sub> ) <sub>2</sub> CO <sub>3</sub>	I	ammonium carbonate

2. Name the following ionic compounds:

a) Na <sub>2</sub> SO <sub>4</sub>	sodium sulfate	b) Ag <sub>2</sub> S	silver sulfide
c) CuNO	copper(I) nitrate	d) Ba <sub>3</sub> (PO <sub>3</sub> ) <sub>2</sub>	barium phosphite

3. Give the formulas for the following ionic compounds:

a) aluminum hydroxide	Al(OH) <sub>3</sub>	b) iron (III) nitride	FeN
c) magnesium phosphate	Mg <sub>3</sub> (PO <sub>4</sub> ) <sub>2</sub>	d) ammonium sulfate	(NH <sub>4</sub> ) <sub>2</sub> SO <sub>4</sub>
e) silver carbonate	Ag <sub>2</sub> CO <sub>3</sub>	f) manganese (II) nitrate	Mn(NO <sub>3</sub> ) <sub>2</sub>

4. Give the formulas for the following covalent compounds:

a) sulfur trioxide	SO <sub>3</sub>	b) hydrobromic acid	HBr
c) carbonic acid	H <sub>2</sub> CO <sub>3</sub>	d) dinitrogen tetrachloride	N <sub>2</sub> Cl <sub>4</sub>